



UNIVERSITÀ DI PARMA

ARCHIVIO DELLA RICERCA

University of Parma Research Repository

Catalytic oxidative carbonylation of amino moieties to ureas, oxamides, 2-oxazolidinones, and benzoxazolones

This is the peer reviewed version of the following article:

Original

Catalytic oxidative carbonylation of amino moieties to ureas, oxamides, 2-oxazolidinones, and benzoxazolones / Mancuso, Raffaella; Raut, Dnyaneshwar S.; DELLA CA', Nicola; Fini, Francesco; Carfagna, Carla; Gabriele, Bartolo. - In: CHEMSUSCHEM. - ISSN 1864-5631. - 8:13(2015), pp. 2204-2211. [10.1002/cssc.201500343]

Availability:

This version is available at: 11381/2797493 since: 2021-09-29T12:58:38Z

Publisher:

Wiley-VCH Verlag

Published

DOI:10.1002/cssc.201500343

Terms of use:

Anyone can freely access the full text of works made available as "Open Access". Works made available

Publisher copyright

note finali coverpage

(Article begins on next page)

Recyclable Catalytic Oxidative Carbonylation of Amines, β -Amino alcohols, and 2-Aminophenols to Ureas, Oxamides, 2-Oxazolidinones, and Benzoxazolones

Raffaella Mancuso,^{*[a,b]} Dnyaneshwar S. Raut,^[a] Nicola Della Ca',^[c] Francesco Fini,^[d] Carla Carfagna,^[e] and Bartolo Gabriele^{*[a]}

Abstract: The direct syntheses of ureas, oxamides, 2-oxazolidinones, and benzoxazolones by oxidative carbonylation of amines, β -amino alcohols, and 2-aminophenols allows obtaining high value added molecules, with a large number of important applications in several fields, starting from very simple building blocks. In this work, we have found that it is possible to carry out these transformations using the PdI₂/KI catalytic system in an ionic liquid (IL), such as BmimBF₄, as the solvent, and that the solvent-catalyst system can be recycled several times with only a slight loss of activity, while the product can be easily recovered by crystallization.

Introduction

Ureas, oxamides, 2-oxazolidinones, and benzoxazolones are between the most important carbonyl derivatives.^[1-11] The classical approaches to these molecules are based on the use of toxic reactants, such as phosgene and its derivatives, or oxalyl chloride and its derivatives.^[12] Recently, the possibility to directly synthesize these important compounds by catalytic oxidative carbonylation of amines or β -amino alcohols has been extensively explored.^[13] In fact, this approach may allow obtaining the target molecules in a multicomponent fashion starting from very simple and largely available building blocks (the organic substrate, CO, and an external oxidant). Moreover, when the oxidant corresponds to molecular oxygen, the process becomes particularly attractive from a sustainable perspective, since the coproduct formed is water and the atom economy is particularly high. Several methods have thus appeared in the

literature for realizing these processes with oxygen as reoxidant.^[14, 15] Unfortunately, most of the proposed methods have been proposed under potentially explosive conditions, with CO/O₂ mixtures in volume ratios very close to, or even within, the flammability range for this kind of mixture.^[14c] Moreover, to the best of our knowledge, no oxidative carbonylation method reported so far has allowed the general synthesis of ureas, oxamides, 2-oxazolidinones, and benzoxazolones with recycling of the catalytic system and of the reaction medium, which would clearly increase the attractiveness of the process and its sustainability.^[16]

Some years ago we found that a very simple catalytic system, consisting of PdI₂ in conjunction with an excess of KI, is an excellent catalyst for performing the oxidative carbonylation of primary amines to 1,2-symmetrically disubstituted ureas (or of a mixture of a primary and a secondary amine to trisubstituted ureas)^[17] as well as of β -amino alcohols bearing a primary amino group to the corresponding oxazolidinones.^[18] Reactions were carried out in 1,2-dimethoxyethane (DME) as the solvent, under non-explosive conditions (typically, 100 °C under 60 atm of a CO/air/CO₂ mixture or 20 atm of a CO/air mixture). The use of an excess of CO₂ in the reaction mixture caused in several cases a significant improvement in the product yield and selectivity. Later on, we also found that, employing the same type of catalytic system, it was possible to carry out the oxidative carbonylation of primary amines to the corresponding ureas in carbon dioxide medium.^[19]

We have now found that by using a suitable ionic liquid (IL),^[20] such as BmimBF₄, as the reaction medium, in the presence of the PdI₂/KI catalytic system, it is possible not only to convert primary aromatic amines to 1,3-aryl ureas, β -amino alcohols to 2-oxazolidinones, and 2-aminophenols to benzoxazolones (through an oxidative *mono* carbonylation process), but also

[a] Dr. Dnyaneshwar S. Raut, Prof. Dr. Bartolo Gabriele
Dipartimento di Chimica e Tecnologie Chimiche
Università della Calabria
Via Pietro Bucci, 12/C, 87036 Arcavacata di Rende (CS), Italy
E-mail: bartolo.gabriele@unical.it

[b] Dr. Raffaella Mancuso
Dipartimento di Ingegneria Meccanica, Energetica e Gestionale
Università della Calabria
87036 Arcavacata di Rende (CS), Italy
E-mail: raffaella.mancuso@unical.it;

[c] Dr. Nicola della Ca'
Dipartimento di Chimica and CIRCC
Università di Parma
Parco Area delle Scienze 17/A, 43124 Parma, Italy

[d] Dr. Francesco Fini
Dipartimento di Scienze Biomolecolari
Università di Urbino
Piazza Rinascimento 6, 61029 Urbino (PU), Italy

[e] Prof. Dr. Carla Carfagna
Dipartimento di Chimica Industriale "Toso Montanari"
Università di Bologna
Viale Risorgimento 4, 40136 Bologna, Italy

Table 1. Synthesis of 1,3-diaryl ureas **2a-e** by recyclable PdI₂/KI-catalyzed oxidative carbonylation of primary anilines **1a-e** in BmimBF₄^[a]

$$2 \text{ ArNH}_2 \text{ (1)} + \text{CO} + (1/2) \text{O}_2 \xrightarrow[\text{-H}_2\text{O}]{\text{PdI}_2/\text{KI}, \text{BmimBF}_4} \text{ArHN} \text{ (2)} \text{C(=O)} \text{NHAr}$$

Entry	Aryl amine 1	Urea 2	Yield of 2 [%] ^[b]					
			Run 1	Run 2	Run 3	Run 4	Run 5	Run 6
1			94	87	85	80	77	75
2			87	81	80	77	75	75
3			92	87	87	85	86	85
4			88	80	78	77	75	75
5			91	80	80	78	77	76

[a] All reactions were carried out at 100 °C for 24 h in BmimBF₄ (0.5 mmol of **1** at start per mL of BmimBF₄, 1 mmol scale based on **1**) using a 1:KI:PdI₂ molar ratio of 100:10:1, under 20 atm (at 25 °C) of a 4:1 mixture of CO-air. Substrate conversion was quantitative. [b] Isolated yield from **1**. Run 1 corresponds to the 1st experiment, the next runs to recycles. See text for details.

secondary amines to tetrasubstituted oxamides (through an oxidative *double* carbonylation process). Moreover, primary aliphatic amines can be selectively converted into either 1,3-dialkyl ureas or *N,N*-dialkyloxalamides depending on reaction conditions. In all cases, the process is carried out under conditions that allow an efficient recovery and recycling of the solvent-catalyst system.

Results and Discussion

Initial experiments were carried out with aniline **1a** as the substrate, using 1 mol % of PdI₂ in conjunction with 10 mol % of KI, at 100 °C under 20 atm of a 4:1 mixture of CO-air, in BmimF₄ as the solvent (0.5 mmol of **1a** per mL of solvent). After 24 h, substrate conversion was quantitative, and extraction with Et₂O followed by repeated crystallization from cold Et₂O allowed the isolation of pure 1,3-diphenylurea **2a** in almost quantitative yield (94%, Table 1, entry 1, run 1). The ether used for extraction and

purification could be easily recovered and reused for other experiments. We then verified the recyclability of the catalyst-ionic liquid system by adding fresh aniline to the recovered ionic liquid phase and repeating the carbonylation procedure under the same conditions reported above, but without adding fresh catalyst. After 24 h, **1a** conversion was still quantitative, and **2a** was recovered in 87% yield (Table 1, entry 1, run 2). The recycling process was repeated for additional 4 times, with formation of **2a** in yields ranging from 75 to 85% (Table 1, entry 1, runs 3-6). These results therefore confirmed the possibility to use an ionic liquid medium in the PdI₂/KI-catalyzed direct oxidative carbonylation of amines to ureas and to recycle the catalyst several times with only a slight loss of activity.

The method was then extended to other primary aromatic amines **1b-e**, and the results obtained are shown in Table 1, entries 2-5. As can be seen from Table 1, excellent results in terms of product yield and catalyst recycling were obtained with

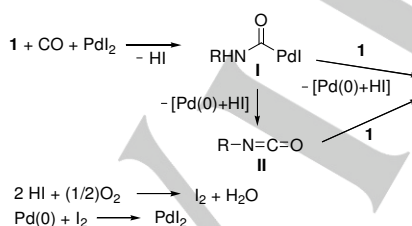
Table 2. Pd₂/KI-catalyzed oxidative carbonylation of primary amines **1f-h** in BmimBF₄ leading to 1,3-dialkylureas **2f-h** and/or *N,N*-dialkylloxalamides **3f-g** [a]

Entry	Alkylamine 1	Dialkylurea 2	Yield of 2 [%] ^[b]	Oxamide 3	Yield of 3 [%] ^[b]
1	BuNH ₂ 1f		40		44
2	<i>t</i> -BuNH ₂ 1g		98 (94-94-93-92-92)		
3	BnNH ₂ 1h		47		45
4 ^[c]	1f	2f	83 (81-80-78-77-77)		
5 ^[d]	1h	2h	90 (89-89-88-88-87)		
6 ^[e]	1f			3f	82 (80-81-79-80-81)
7 ^[e]	1h			3h	86 (85-85-84-83-84)

[a] Unless otherwise noted, all reactions were carried out at 100 °C for 24 h in BmimBF₄ (0.5 mmol of **1** at start per mL of BmimBF₄, 1 mmol scale based on **1**) using a **1**:KI:Pd₂ molar ratio of 100:10:1, under 20 atm (at 25 °C) of a 4:1 mixture of CO-air. Substrate conversion was quantitative. [b] Isolated yields from **1**. Figures in parentheses refer to recycles. [c] The reaction was carried out with a **1f**:KI:Pd₂ molar ratio of 200:10:1 and a substrate concentration of 1 mmol of **1f** per mL of BmimBF₄. [d] The reaction was carried out with a **1h**:KI:Pd₂ molar ratio of 500:10:1 and a substrate concentration of 1 mmol of **1h** per mL of BmimBF₄. [e] The reaction was carried out at 80 °C under 40 atm (at 25 °C) of a 4:1 mixture of CO-air.

all the amines tested, bearing electron-donating (Table 1, entries 2-4) as well as electron-withdrawing (Table 1, entry 5) groups.

Mechanistically, the reaction may start with the formation of a carbamoylpalladium iodide intermediate **I**, from the reaction between PdI₂, the primary amine **1**, and CO (Scheme 1; anionic iodide ligands are omitted for clarity).^[14c, 17-19]

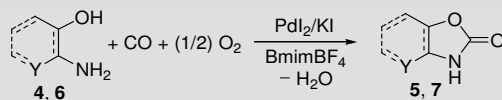


Scheme 1. Possible mechanistic pathways for the Pd₂-catalyzed oxidative monocarbonylation of primary amines **1a-h** to ureas **2a-h**. Both carbamoylpalladium complex **I** and isocyanate intermediate **II** may lead to the final product.

Either the direct nucleophilic displacement by the amine, or the formation of an isocyanate intermediate **II** (by formal β-

elimination of HI from the I-Pd-(CO)NHR moiety) followed by addition of **1** to **II**, may then occur, leading to the formation of the final urea product **2** and Pd(0). The latter is eventually reoxidized back to PdI₂, according to the mechanism we demonstrated several years ago,^[21] involving oxidation of HI (formally deriving from the carbonylation process) to I₂ followed by oxidative addition of I₂ to Pd(0).

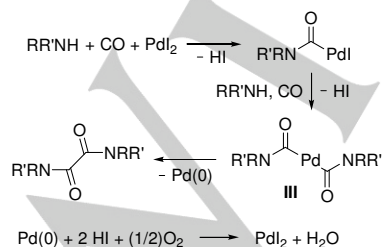
When butylamine **1f** was subjected to the same carbonylation conditions as those reported in Table 1, a mixture of 1,3-dibutylurea **2f** and *N,N*-dibutylloxalamide **3f** was obtained, as confirmed by GC-MS analysis (total yield: 84%; **3f/2f** molar ratio = 1.1, Table 2, entry 1). The formation of **3f**, ensuing from a double carbonylation process, is clearly due to the higher nucleophilicity of **1f** as compared to anilines **1a-e**, which allows for the formation of a *bis*-carbamoylpalladium complex Pd(CONHR)₂ **III** (from the reaction between PdI₂ and 2 mol of RNH₂ and CO, with formal elimination of 2 HI; Scheme 2, R = alkyl, R' = H). The latter, according to the literature,^[14] is the key intermediate in the formation of the oxamide product, obtained by reductive elimination (Scheme 2). Accordingly, a less nucleophilic, sterically hindered primary aliphatic amine, such as *tert*-butylamine **1g**, selectively led to the urea derivative in practically quantitative yield (Table 2, entry 2), while

Table 3. Synthesis of 2-oxazolidinones **5** and benzoxazolones **7** by recyclable Pd₂/KI-catalyzed oxidative carbonylation of β-amino alcohols **4** and 2-aminophenols **6**, respectively, in BmimBF₄^[a]

Entry	β-Aminoalcohol 4 or 2-aminophenol 6	2-Oxazolidinone 5 or benzoxazolone 7	Yield of 5 or 7 [%] ^[b]					
			Run 1	Run 2	Run 3	Run 4	Run 5	Run 6
1			94	88	86	85	83	84
2			80	75	74	72	73	72
3			98	93	90	88	85	85
4 ^[c]	6a	7a	95	94	90	88	88	89
5			92	90	88	86	85	86
6			90	87	85	83	84	84

[a] Unless otherwise noted, all reactions were carried out at 100 °C in BmimBF₄ (0.5 mmol of **5** at start per mL of BmimBF₄, 1 mmol scale based on **5**) using a 5:KI:Pd₂ molar ratio of 100:10:1, under 20 atm (at 25 °C) of a 4:1 mixture of CO-air. Substrate conversion was quantitative. [b] Isolated yield from **5**. Run 1 corresponds to the 1st experiment, the next runs to recycles. [c] The reaction was carried out with a **5a**:KI:Pd₂ molar ratio of 500:10:1.

benzylamine **1h** behaved similarly to butylamine **1f** (Table 2, entry 3).

**Scheme 2.** Formation of oxamides by Pd₂-catalyzed oxidative double carbonylation of amines through the formation of *bis*-carbamoylpalladium complex **III**.


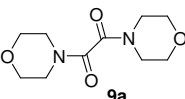

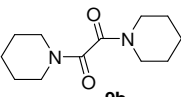
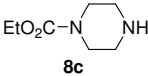
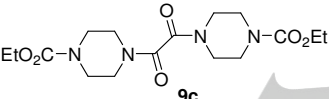

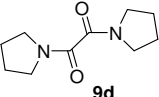
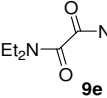
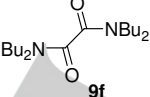
Very interestingly, both for BuNH₂ **1f** and BnNH₂ **1h**, it has been possible to work out suitable reaction conditions able to selectively promote the formation of either the urea or the oxamide derivative. In particular, when the carbonylation reaction of **1f** was carried out with a higher substrate concentration (1 mmol per mL of BmimBF₄) with a catalyst loading of 0.5%, urea **2f** was the only product formed in 83% yield (Table 2, entry 4). Under similar conditions, and with a catalyst loading as low as 0.2%, benzylamine **1h** was converted into dibenzylurea **2h** in 90% yield (Table 2, entry 5). On the other hand, oxamides **3f** and **3h** were selectively formed when the reactions were conducted at 80 °C under a total pressure of 40 atm (entries 6 and 7 of Table 2). As shown in Table 2, entries 2 and 4-7, the recyclability of the catalyst was successfully verified also in these experiments.

In the case of β-amino alcohols **4** as starting materials, an intramolecular nucleophilic attack may take place at level of the

Table 4. Synthesis of tetrasubstituted oxamides **9** by recyclable Pd₂/KI-catalyzed oxidative carbonylation of secondary amines **8** in BmimBF₄^[a]

$$2 \text{ R}_2\text{NH} + 2 \text{ CO} + (1/2) \text{ O}_2 \xrightarrow[\text{-H}_2\text{O}]{\text{Pd}_2/\text{KI}, \text{BmimBF}_4} \text{R}_2\text{N}-\text{C}(=\text{O})-\text{NR}_2$$

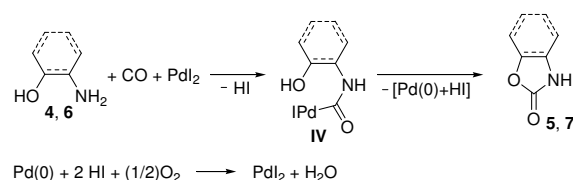
8 **9**

Entry	Dialkyl amine 8	Oxamide 9	Yields of 9 [%] ^[b]					
			Run 1	Run 2	Run 3	Run 4	Run 5	Run 6
1			85	86	85	85	85	87
2 ^[c]	8a	9a	85	83	83	90	81	81
3			91	90	92	91	90	91
4			97	95	96	94	94	92
5			90	90	91	90	89	91
6	Et ₂ NH 8e		92	90	87	85	85	86
7	Bu ₂ NH 8f		89	86	84	83	83	82

[a] Unless otherwise noted, all reactions were carried out at 100 °C in BmimBF₄ (0.5 mmol of **8** at start per mL of BmimBF₄, 1 mmol scale based on **8**) using a **8**:KI:Pd₂ molar ratio of 100:10:1, under 20 atm (at 25 °C) of a 4:1 mixture of CO-air. Substrate conversion was quantitative. [b] Isolated yield from **8**. Run 1 corresponds to the 1st experiment, the next runs to recycles. [c] The reaction was carried out with a **8a**:KI:Pd₂ molar ratio of 500:10:1.

β-hydroxycarbonylpalladium complex **IV**, with formation of particularly important products, 2-oxazolidinones (Scheme 3).

Accordingly, the reactivity of 2-aminopropan-1-ol **4a** and (*R*)-2-amino-2-phenylethanol **4b** was tested under our reaction conditions. Excellent yields in the corresponding 4-methyloxazolidin-2-one **5a** and (*R*)-4-phenyloxazolidin-2-one **5b** were obtained (Table 3, entries 1 and 2). Again, 2-oxazolidinones were recovered by simple crystallization, while the ionic liquid-catalyst system could be successfully recycled several times.

**Scheme 3.** Formation of 2-oxazolidinones **5** and benzoxazolones **7** by Pd₂-catalyzed oxidative monocarbonylation of β-amino alcohols **4** or 2-aminophenols **6**, respectively, by intramolecular trapping of the β-hydroxycarbonylpalladium complex **IV**.

Extension to 2-aminophenols **6a,b** to give the corresponding benzoxazolones **7a,b** was also successful (Table 3, entries 3-5), and the possibility to lower the catalyst loading to 0.2 mol% verified (Table 3, entry 4). A pyridine ring was compatible with the reactions conditions, 2-aminopyridin-3-ol **6c** being converted into oxazolo[4,5-*b*]pyridin-2(3*H*)-one **7c** in 90% yield in the first experiment, and in 84-87% yield in the recycles (Table 3, entry 6).

In sharp contrast to our previously reported method,^[17] involving the use of DME as the solvent, which did not work with secondary amines **8**, this class of compounds was smoothly converted into tetrasubstituted oxamides **9** under the conditions reported in the present work, through the intermediate formation of the corresponding *bis*-carbamoylpalladium complex Pd(CONRR')₂ **III** (Scheme 2, R, R' = alkyl). Thus, the reaction of morpholine **8a**, carried out in BmimBF₄ under the same conditions reported in Table 1, entry 1, run 1, led to the formation of 1,2-dimorpholinoethane-1,2-dione **9a** in 85% isolated yield, after extraction of the reaction mixture with Et₂O (Table 4, entry 1).^[22] Also in this case, the ethereal solvent used for crystallization could be recovered and reused, while the IL phase, still containing the active catalyst, could be recycled several times using a fresh amount of substrate **8a** in each run, without appreciable loss of activity (Table 4, entry 1, runs 2-6). The oxidative double carbonylation reaction leading to **9a** worked nicely even when the catalyst loading was lowered to 0.2% (Table 4, entry 2). Under the same conditions employed for morpholine **8a**, other cyclic and acyclic secondary amines **8b-f** were efficiently converted into the corresponding oxamides **9b-f** in high yields and excellent recyclability of the catalyst-IL system (Table 4, entries 2-7).

Conclusions

In conclusion, we have reported the first general recyclable Pd-catalyzed oxidative carbonylation of primary amines, β-amino alcohols, 2-aminophenols, and secondary amines to give 1,3-disubstituted ureas, 2-oxazolidinones, benzoxazolones, and oxamides, in high yields and selectivity. Carbonylations were carried out in an ionic liquid, such as BmimBF₄, as the reaction medium, under non-explosive conditions (100 °C and 20 atm of a 4:1 mixture of CO-air), using a particularly simple and robust catalytic system, consisting of PdI₂ in conjunction with KI. While primary aromatic amines selectively led to 1,3-aryl ureas, tetrasubstituted oxamides were formed from secondary amines. On the other hand, primary aliphatic amines could be selectively converted into either 1,3-dialkyl ureas or *N,N*-dialkyloxalamides depending on reaction conditions. Oxazolidinones and benzoxazolones were obtained from β-amino alcohols and 2-aminophenols, respectively. All products were easily recovered by simple extraction with Et₂O from the reaction mixture followed by crystallization. The ethereal solvent was recovered and reused, while the IL phase, still containing the active catalyst dissolved in it, could be recycled several times without appreciable loss of activity.

Experimental Section

All solvents and chemicals were purchased from Sigma-Aldrich, were reagent grade and used without further purification. All reactions were analyzed by TLC on silica gel 60 F254 (Merck), or by GLC, using a Shimadzu GC-2010 gas chromatograph using capillary columns with polymethylsilicone+5% phenylsilicone as the stationary phase (HP-5). Column chromatography was performed on silica gel 60 (Merck, 70–230 mesh). Evaporation refers to the removal of solvent under reduced pressure.

Melting points are uncorrected. ¹H and ¹³C NMR spectra were recorded at 25 °C on a Bruker DPX Avance 300 Spectrometer at 300 and 75 MHz, respectively, with Me₄Si as internal standard. IR spectra were recorded on a JASCO FT-IR 4200 spectrometer. GC-MS spectra were recorded on a Shimadzu QP-2010 GC-MS apparatus at 70 eV ionization voltage. The electrospray ionization mass spectra were acquired by direct infusion on MSD Sciex Applied Biosystem API 4000 Q-Trap mass spectrometer equipped with a turbo ion spray ionization source. The spectra in the positive ion mode were obtained at ionspray voltage 5500 V, curtain gas 20 psi, temperature 25 °C, ion source gas 12 psi, collision gas thickness (CAD) medium; entrance potential (EP) between 6 and 10 eV; declustering potential (DP) between 70 and 85 eV; entrance collision energy (CE) between 15 and 30 eV and exit collision energy (CXP) between 8 and 10 eV. Microanalyses were carried out with a Thermo-Fischer Elemental Analyzer Flash 2000.

All starting materials (amines **1a-h** and **8a-8f**, β-amino alcohols **4a-b**, 2-aminophenols **6a-c**) were commercially available (Sigma-Aldrich) and were used as received.

Characterization data for all products are given in the Supporting Information.

Recyclable synthesis of 1,3-disubstituted ureas **2a-h** (Tables 1 and 2)

The general procedure for the recyclable catalyzed oxidative carbonylation of primary amines **1a-h** is as follows: A 100 mL stainless steel autoclave was charged in air with PdI₂ (4.0 mg, 1.11×10⁻² mmol), KI (18.4 mg, 1.11×10⁻¹ mmol), and a solution of **1** [**1a**, 103 mg (1.11 mmol, Table 1, entry 1); **1b**, 120 mg (1.12 mmol, Table 1, entry 5); **1c**, 150 mg (1.11 mmol, Table 1, entry 3); **1d**, 137mg (1.11 mmol, Table 1, entry 4); **1e**, 142 mg (1.11 mmol, Table 1, entry 5); **1f**, 162 mg (2.22 mmol, Table 2, entry 4); **1g**, 81 mg (1.11 mmol, Table 2, entry 2); **1h**, 595 mg (5.55 mmol Table 2, entry 5)] in BmimBF₄ (for **1a-g**: 2.2 mL; for **1h**: 5.5 mL); see Tables 1, 2. The autoclave was sealed and, with stirring of the mixture, pressurized with CO (16 atm) and air (4 atm). After stirring at 100 °C for 24 h, the autoclave was cooled, degassed, and opened. The crude reaction mixture was transferred to a round-bottom flask, and Et₂O (for **1a**, **1c-h**: 5 mL; for **1b**: 10 mL) was added. The mixture was stirred for 15 min, and the ethereal phase was decanted. Part of the urea product was thus extracted in Et₂O, while some part remained in the IL phase as a white solid in suspension. To recover the product suspended in the IL phase, CH₂Cl₂ (for **1a**, **1c-h**: 3 mL; for **1b**: 8 mL) was added to the IL phase, followed by stirring for 5 min. Excess of CH₂Cl₂ was evaporated, the distilled CH₂Cl₂ recovered in order to be reused again, and the IL residue re-extracted with Et₂O (for **1a**, **1c-h**: 5 mL; for **1b**: 10 mL). The extraction process was repeated for additional five times. The collected ethereal phases were evaporated to remove about the 80% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Crystals of the pure urea product separated, and were collected by decantation (1st

crop). The decanted solution was evaporated to remove about the 50% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Additional crystals of the pure urea product separated were collected by decantation (2nd crop). The collected crystals (1st crop + 2nd crop) were eventually dried in vacuo. **1,3-Diphenylurea (2a)**: 111 mg yield (94% from **1a**; Table 1, entry 1, run 1); **1,3-di-*p*-tolylurea (2b)**: 116 mg yield (87% from **1b**; Table 1, entry 2, run 1); **1,3-bis-(4-isopropylphenyl)urea (2c)**: 151 mg yield (92% from **1c**; Table 1, entry 3, run 1); **1,3-bis-(4-methoxyphenyl)urea (2d)**: 133 mg yield (88% from **1d**; Table 1, entry 4, run 1); **1,3-bis-(4-chlorophenyl)urea (2e)**: 142 mg yield (91% from **1e**; Table 1, entry 5, run 1); **1,3-dibutylurea (2f)**: 159 mg yield (83% from **1f**; Table 2, entry 4, run 1); **1,3-di-*tert*-butylurea (2g)**: 94 mg yield (98% from **1g**; Table 2, entry 2, run 1); **1,3-dibenzylurea (2h)**: 600 mg yield (90% from **1h**; Table 2, entry 5, run 1).

Recycling Procedure: The IL residue obtained as described above, still containing the catalytic system dissolved in it, was dried in vacuo for 15 h, in order to remove traces of Et₂O, CH₂Cl₂, and moisture. The dried IL residue was then used as such for the next recycle. Thus, it was transferred into the 100 mL stainless steel autoclave, the amine substrate was added, and then the same procedure described above was followed.

Recyclable synthesis of disubstituted oxamides **3f** and **3h** (Table 2)

The general procedure for the recyclable catalyzed oxidative carbonylation of primary aliphatic amines **1f** and **1h** is as follows: A 100 mL stainless steel autoclave was charged in air with PdI₂ (4.0 mg, 1.11×10⁻² mmol), KI (18.4 mg, 1.11×10⁻¹ mmol), and a solution of **1** [(1.11 mmol); **1f**, 81 mg (Table 2, entry 6); **1h**, 119 mg (Table 2, entry 7)] in BmimBF₄ (2.2 mL); see Table 2. The autoclave was sealed and, with stirring of the mixture, pressurized with CO (32 atm) and air (8 atm). After stirring at 80 °C for 24 h, the autoclave was cooled, degassed, and opened. The crude reaction mixture was transferred to a round-bottom flask, and Et₂O (5 mL) was added. The mixture was stirred for 15 min, the ethereal phase was decanted, and the IL residue was re-extracted with Et₂O (5 × 5 mL). The collected ethereal phases were evaporated to remove about the 80% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Crystals of the pure product separated, and were collected by decantation (1st crop). The decanted solution was evaporated to remove about the 50% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Additional crystals of the pure product separated, and were collected by decantation (2nd crop). The collected crystals (1st crop + 2nd crop) were eventually dried in vacuo. *N,N*-Dibutylalamide (**3f**): 91 mg yield (82% from **1f**; Table 2, entry 6, run 1); *N,N*-dibenzylalamide (**3h**): 128 mg yield (86% from **1h**; Table 2, entry 7, run 1).

Recycling Procedure: The IL residue obtained as described above, still containing the catalytic system dissolved in it, was dried in vacuo for 15 h, in order to remove traces of Et₂O, CH₂Cl₂, and moisture. The dried IL residue was then used as such for the next recycle. Thus, it was transferred into the 100 mL stainless steel autoclave, the amine substrate was added, and then the same procedure described above was followed.

Recyclable synthesis of 2-oxazolidinones **5a** and **5b** and benzoxazolones **7a-c** (Table 3)

The general procedure for the recyclable catalyzed oxidative carbonylation of β-amino alcohols **4a** and **4b** and 2-aminophenols **6a-c** is

as follows. A 100 mL stainless steel autoclave was charged in air with PdI₂ (4.0 mg, 1.11×10⁻² mmol), KI (18.4 mg, 1.11×10⁻¹ mmol), and a solution of **4** [1.11 mmol; **4a**, 83 mg (Table 3, entry 1); **4b**, 152 mg (Table 3, entry 2)] or **6** [1.11 mmol; **6a**, 121 mg (Table 3, entry 3); **6b**, 137 mg (Table 3, entry 5); **6c**, 122 mg (Table 2, entry 6)] in BmimBF₄ (2.2 mL); see Table 3. The autoclave was sealed and, with stirring of the mixture, pressurized with CO (16 atm) and air (4 atm). After stirring at 100 °C for 24 h, the autoclave was cooled, degassed, and opened. The crude reaction mixture was transferred to a round-bottomed flask, and Et₂O (5 mL) was added. The mixture was stirred for 15 min, and the ethereal phase was decanted, and the IL residue was re-extracted with Et₂O (5 × 5 mL).

In the case of the reaction of **4a**, the collected ethereal phases were evaporated, and the residue purified by column chromatography on silica gel using 95:5 hexane/AcOEt to give pure 4-methylloxazolidin-2-one **5a** (105 mg yield; 94% from **4a**; Table 3, entry 1, run 1).

In the case of the reactions of **4b** and **6a-c**, the collected ethereal phases were evaporated to remove about the 80% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Crystals of the pure product separated, and were collected by decantation (1st crop). The decanted solution was evaporated to remove about the 50% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Additional crystals of the pure product separated, and were collected by decantation (2nd crop). The collected crystals (1st crop + 2nd crop) were eventually dried in vacuo. (*R*)-4-Phenylloxazolidin-2-one (**5b**): 144 mg yield (80% from **4b**; Table 3, entry 2, run 1); benzo[*d*]oxazol-2(3*H*)-one (**7a**): 147 mg yield (98% from **6a**; Table 3, entry 3, run 1); 6-methylbenzo[*d*]oxazol-2(3*H*)-one (**7b**): 153 mg yield (92% from **6b**; Table 3, entry 5, run 1); oxazolo[4,5-*b*]pyridin-2(3*H*)-one (**7c**): 136 mg yield (90% from **6c**; Table 3, entry 6, run 1).

With 2-aminophenol **6a**, the oxidative carbonylation reaction was also carried out using only 0.2 mol % of catalyst (Table 3, entry 4). The general procedure described above was followed, starting from 606 mg of **6a** (5.55 mmol) in 11.1 mL of BmimBF₄. The reaction mixture was then extracted with 6 × 10 mL of Et₂O, and the product separated by crystallization as described above (yield: 710 mg, 95% from **6a**; Table 3, entry 4, run 1).

Recycling Procedure: The IL residue obtained as described above, still containing the catalytic system dissolved in it, was dried in vacuo for 15 h, in order to remove traces of Et₂O and moisture. The dried IL residue was then used as such for the next recycle. Thus, it was transferred into the 100 mL stainless steel autoclave, the substrate was added, and then the same procedure described above was followed.

Recyclable synthesis of tetrasubstituted oxamides **9a-f** (Table 4)

The general procedure for the recyclable catalyzed oxidative carbonylation of secondary amines **8a-f** is as follows. A 100 mL stainless steel autoclave was charged in air with PdI₂ (4.0 mg, 1.11×10⁻² mmol), KI (18.4 mg, 1.11×10⁻¹ mmol), and a solution of **8** [1.11 mmol; **8a**, 97 mg (Table 4, entry 1); **8b**, 95 mg (Table 4, entry 3); **8c**, 176 mg (Table 4, entry 4); **8d**, 75 mg (Table 4, entry 5); **8e**, 81 mg (Table 4, entry 6), **8f**, 112 mg (Table 4, entry 7)] in BmimBF₄ (2.2 mL); see Table 4. The autoclave was sealed and, with stirring of the mixture, pressurized with CO (16 atm) and air (4 atm). After stirring at 100 °C for 24 h, the autoclave was cooled, degassed, and opened. The crude reaction

mixture was transferred to a round-bottom flask, and Et₂O (5 mL) was added. The mixture was stirred for 15 min, the ethereal phase was decanted, and the IL residue was re-extracted with Et₂O (5 × 5 mL).

In the case of the reactions of **8a-e**, the collected ethereal phases were evaporated to remove about the 80% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Crystals of the pure product separated, and were collected by decantation (1st crop). The decanted solution was evaporated to remove about the 50% of the solvent, and the distilled Et₂O recovered in order to be reused again. The residual ethereal solution was kept at 0 °C for 15 h. Additional crystals of the pure product separated, and were collected by decantation (2nd crop). The collected crystals (1st crop + 2nd crop) were eventually dried in vacuo. **1,2-Dimorpholinoethane-1,2-dione (9a)**: 108 mg yield (85% from **8a**; Table 4, entry 1, run 1); **1,2-di(piperidin-1-yl)ethane-1,2-dione (9b)**: 113 mg yield (91% from **8b**; Table 4, entry 3, run 1); **1,2-di-(4-ethoxycarbonyl)piperazinoethane-1,2-dione (9c)**: 200 mg yield (97% from **8c**; Table 4, entry 4, run 1); **1,2-di(pyrrolidin-1-yl)ethane-1,2-dione (9d)**: 98 mg yield (90% from **8d**; Table 4, entry 5, run 1); **tetraethyloxalamide (9e)**: 102 mg yield (92% from **8e**; Table 4, entry 6, run 1).

In the case of the reaction of **8f**, the collected ethereal phases were evaporated, and the residue purified by column chromatography on silica gel using 8:2 hexane/AcOEt to give pure tetraisopropyloxalamide **9f** (Yield: 154 mg; 89% from **8f**; Table 4, entry 7, run 1).

With morpholine **8a**, the oxidative carbonylation reaction was also carried out using only 0.2 mol % of catalyst (Table 4, entry 2). The general procedure described above was followed, starting from 484 mg of **8a** (5.55 mmol) in 11.1 mL of BmimBF₄. The reaction mixture was then extracted with 6 × 10 mL of Et₂O, and the product separated by crystallization as described above. (Yield: 540 mg; 85% from **8a**; Table 4, entry 2, run 1).

Recycling Procedure: The IL residue obtained as described above, still containing the catalytic system dissolved in it, was dried in vacuo for 15 h, in order to remove traces of Et₂O and moisture. The dried IL residue was then used as such for the next recycle. Thus, it was transferred into the 100 mL stainless steel autoclave, the substrate was added, and then the same procedure described above was followed.

Acknowledgements

This research work was partially supported by MIUR PON R&C 2007-2013 Project "DIRECT FOOD" cod. PON01_00878. We also thank Dr. Enzo Perri and Dr. Cinzia Benincasa (Centro di ricerca per l'olivicultura e l'industria olearia, Rende, CS, Italy) for LC-MS measurements.

Keywords: carbonylation • ureas • oxamides • oxazolidinones • benzoxazolones • palladium • recyclable synthesis • ionic liquids

[1] For applications of ureas as gelators, see: a) M. Yamanaka, *J. Incl. Phenom. Macrocycl. Chem.* **2013**, *77*, 33-48. For applications of ureas as reagents and catalysts in organic synthesis, see: b) N. Volz, J. Clayden, *Angew. Chem. Int. Ed.* **2011**, *50*, 12148-12155.

- [2] For applications of ureas as receptors for anion recognition, see: V. Amendola, L. Fabbri, L. Mosca, *Chem. Soc. Rev.* **2010**, *39*, 3889-3915.
- [3] For applications of ureas as anti-cancer agents, see: H.-Q. Li, P.-C. Lv, T. Yan, H.-L. Zhu, *Anti-Cancer Agents Med. Chem.* **2009**, *9*, 471-480.
- [4] For applications of ureas as regulators of cell division and differentiation in plants, see: A. Ricci, C. Bertolotti, *Plant Biol.* **2009**, *11*, 262-272.
- [5] For applications of ureas as biosensors, see: a) G. Dhawan, G. Sumana, B. D. Mhotra, *Biochem. Eng. J.* **2009**, *44*, 42-52; b) M. Singh, N. Verma, A. K. Garg, N. Redhu, *Sens. Actuators B-Chem.* **2008**, *134*, 345-351.
- [6] For applications of ureas as protein kinase inhibitors, see: J. Dumas, R. A. Smith, T. B. Lowinger, *Curr. Opin. Drug Discov. Dev.* **2004**, *7*, 600-616.
- [7] For applications of ureas as HIV protease inhibitors, see: T. Yoshida, T. Fujita, H. Chuman, *Curr. Comput.-Aided Drug Des.* **2009**, *5*, 38-55.
- [8] For applications of oxamides as ligands in coordination chemistry, see: a) G. G. Briand, N. Burford, *Adv. Inorg. Chem.* **2000**, *50*, 285-357. b) R. Ruiz, J. Fausm F. Lloret, M. Julve, Y. Journaux, *Coord. Chem. Rev.* **1999**, *193*, 1069-1117. c) H. Ojima, K. Nonoyama, *Coord. Chem. Rev.* **1988**, *92*, 85-111. For applications of oxamides as gelators, see: d) L. Frkanec, M. Zinic, *Chem. Commun.* **2010**, *46*, 522-537; e) F. Fages, F. Vogtle, M. Zinic, *Top. Curr. Chem.* **2005**, *256*, 77-131.
- [9] For reviews on the applications of oxazolidinones as antimicrobial agents, see: a) K. Michalska, I. Karpiuk, M. Krol, S. Tyski, *Bioorg. Med. Chem.* **2013**, *21*, 577-591; b) N. Chandna, J. K. Kapoor, V. Goyal, N. K. Aggarwal, K. M. Kumari, M. Vijulatha, *Curr. Top. Med. Chem.* **2013**, *13*, 2062-2075; c) A. R. Renslo, *Expert Rev. Anti-Infect. Ther.* **2010**, *8*, 565-574; d) G. Poce, G. Zappia, G. C. Porretta, B. Botta, M. Biava, *Expert Opin. Ther. Patents* **2008**, *18*, 97-221; e) G. Zappia, P. Menendez, G. Delle Monache, D. Misiti, L. Nevola, B. Botta, *Mini-Rev. Med. Chem.* **2007**, *7*, 389-409; f) J. Vara Prasad, *Curr. Opin. Microbiol.* **2007**, *10*, 454-460; g) A. R. Renslo, G. W. Luehr, M. F. Gordeev, *Bioorg. Med. Chem.* **2006**, *14*, 4227-4240; h) M. B. Gravestock, *Curr. Opin. Drug Discov. Develop.* **2005**, *8*, 469-477; i) T. A. Mukhtar, G. D. Wright, *Chem. Rev.* **2005**, *105*, 529-542; j) D. K. Hutchinson, *Expert Opin. Ther. Patents* **2005**, *14*, 1309-1328; k) B. Bozdogan, P. C. Appelbaum, *Int. J. Antimicrob. Agents* **2004**, *23*, 113-119; l) M. R. Barbachyn, C. W. Ford, *Angew. Chem. Int. Ed.* **2003**, *42*, 2010-2023; m) D. K. Hutchinson, *Curr. Top. Med. Chem.* **2003**, *3*, 1021-1042; n) D. J. Diekema, R. N. Jones, *Lancet* **2001**, *358*, 1975-1982; o) M. J. Genin, *Expert Opin. Ther. Patents* **2000**, *10*, 1405-1414; p) D. J. Diekema, R. N. Jones, *Drugs* **2000**, *59*, 7-16; q) S. M. Abdel-Rahman, G. L. Kearns, *Pediatr. Infect. Dis. J.* **2000**, *18*, 1101-1102; r) M. Muller, K. L. Schimz, *Cell. Mol. Life Sci.* **1999**, *56*, 280-285; s) S. J. Brickner, *Curr. Pharm. Design* **1996**, *2*, 175-194.
- [10] For recent examples of bioactive 2-benzoxazolones, see: a) S. H. Bae, M. J. Kwon, J. B. Park, D. Kim, D.-H. Kim, J.-S. Kang, C.-G. Kim, E. Oh, S. K. Bae, *Antimicrob. Agents Chemother.* **2014**, *58*, 5036-5046; b) N. Ullah, *J. Enzym. Inhib. Med. Chem.* **2014**, *29*, 281 - 291; c) D. Maag, C. Dalvit, D. Thevenet, A. Koehler, F. C. Wouters, D. G. Vassao, J. Gershenzon, J.-L. Wolfender, T. C. J. Turlings, M. Erb, G. Glauser, *Phytochemistry* **2014**, *102*, 97 - 105. d) Z. Zheng, J. B. Parson, R. Tangallapally, W. Zhang, C. O. Rock, R. E. Lee, *Bioorg. Med. Chem. Lett.* **2014**, *24*, 2585 - 2588; e) L. Salerno, V. Pittala, M. N. Modica, M. A. Siracusa, S. Intagliata, A. Cagnotto, M. Salmona, R. Kurczab, A. J. Andrzej, G. Romeo, *Eur. J. Med. Chem.* **2014**, *85*, 716 - 726; f) L. Tang, W.-H. Ma, Y.-L. Ma, S.-R. Ban, X.-E. Feng, Q.-S. Li, *Bioorg. Med. Chem. Lett.* **2013**, *23*, 2380-2383; g) S. Haider, M. S. Alam, H. Hamid, S. Shafi, A. Nargotra, P. Mahajan, S. Nazreen, A. M. Kalle, C. Kharbada, Y. Ali, A. ALam, A. K. Panda, *Eur. J. Med. Chem.* **2013**, *70*, 579 - 588; h) V. E. V. Ferrero, G. Di Nardo, G. Catucci, S. J. Sadeghi, G. Gilardi, *Dalton Trans.* **2012**, *41*, 2018-2025; i) M. Karlgren, A. Vildhede, U. Norinder, J. R. Wisniewski, E. Kimoto, Y. Lai, U. Haglund, P. Artursson, *J. Med. Chem.* **2012**, *55*, 4740-4764.

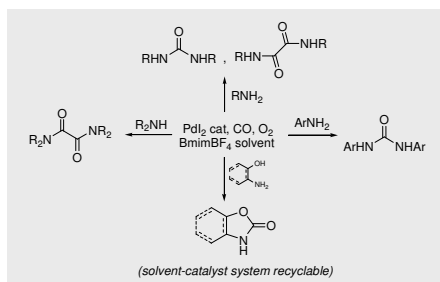
- [11] For a review on the capacity of 2-benzoxazolones to mimic a phenol or a catechol moiety, see: J. Poupaert, P. Carato, E. Colacillo, *Curr. Med. Chem.* **2005**, *12*, 877-885.
- [12] It should be noted, however, that for small products, phosgene may be employed because it can be used in multi-purpose plants that work at low pressure and with a minimized phosgene hold-up. We thank a referee for drawing our attention on this point.
- [13] Oxidative carbonylations may be commercially important, as shown by the Enichem process for the Cu-catalyzed oxidative carbonylation of methanol to dimethyl carbonate: N. Di Muzio, C. Fusi, F. Rivelli, G. Sasselli (Enichem Sintesi, Italy), US Pat. US5210269A (1993).
- [14] For reviews, see: a) Gadge, S. T.; Bhanage, B. M. *RSC Adv.* **2014**, *4*, 10367-10389; b) Wu, X.-F.; Neumann, H.; Beller, M. *ChemSusChem* **2013**, *6*, 229-241; c) B. Gabriele, R. Mancuso, G. Salerno, *Eur. J. Org. Chem.* **2012**, 6825-6839; d) D. J. Díaz, A. K. Darko, L. McElwee-White, *Eur. J. Org. Chem.* **2007**, 4453-4465; e) B. Gabriele, G. Salerno, M. Costa, *Top. Organomet. Chem.* **2006**, *18*, 239-272. For very recent examples, see: f) A. Noujima, T. Mitsudome, T. Mizugaki, K. Jitsukawa, K. Kaneda, *Green Chem.* **2013**, *15*, 608-611 (Au-catalyzed oxidative carbonylation of 2-aminophenols to 2-benzoxazolones); g) T. Mitsudome, A. Noujima, T. Mizugaki, J. Koichiro, *Chem. Commun.* **2012**, *48*, 11733-11735 (Au-catalyzed oxidative carbonylation of amines to oxamides); h) C. Mollar, C. Ramirez de Arellano, M. M. Simón, G. Asensio, *J. Org. Chem.* **2012**, *77*, 9693-9701 (oxidative carbonylation of amines to ureas or oxamides and β -amino alcohols to 2-oxazolidinones).
- [15] For recent patent literature, see: a) F. Gaertner, A. Jacob, J. Sundermeyer, S. Klein, S. Wershofen (Bayer IP GmbH, Germany), Eur. Pat. EP2603319 (A1) (2013); b) F. Gaertner, A. Jacob, J. Sundermeyer, S. Klein, S. Wershofen (Bayer Materials Science AG, F. Gaertner, A. Jacob, J. Sundermeyer, S. Klein, S. Wershofen, Germany), World Pat. WO2012020028 A(1) (2012). It should be noted that the method described in these patents works under potentially explosive CO-O₂ mixtures (36/4 atm at 165 °C).
- [16] The possibility to recycle a heterogeneous Au-based catalysts in the oxidative carbonylation of 2-aminophenols to benzoxazolones has been recently reported: see reference [14f]. Also, very recently, a heterogeneous palladium catalyst has been used for the oxidative carbonylation of amines to ureas and oxamates: S. T. Gadge, E. N. Kusumawati, K. Harada, T. Sasaki, D. Nishio-Hamane, B. M. Bhanage, *J. Mol. Catal. A-Chem.* **2015**, *400*, 170-178; however, the operating conditions (CO/O₂ = 6:1 at 60 °C) were within the flammability range for a CO/O₂ mixture (the flammability range for CO in O₂ is 16.7-93.5% at room temperature, and it becomes even larger at higher temperatures. See: C. M. Bartish, G. M. Drissel, in: *Kirk-Othmer Encyclopedia of Chemical Technology*, vol. 4., 3rd ed. (Eds.: M. Grayson, D. Eckroth, G. J. Bushey, L. Campbell, A. Klingsberg, L. nan Nes), John Wiley & Sons, New York, **1978**, p. 774).
- [17] a) B. Gabriele, R. Mancuso, G. Salerno, M. Costa, *Chem. Commun.* **2003**, 486-487; b) B. Gabriele, G. Salerno, R. Mancuso, M. Costa, *J. Org. Chem.* **2004**, *69*, 4741-4750.
- [18] a) B. Gabriele, G. Salerno, D. Brindisi, M. Costa, G. P. Chiusoli, *Org. Lett.* **2000**, *2*, 625-627; b) B. Gabriele, R. Mancuso, G. Salerno, M. Costa, *J. Org. Chem.* **2003**, *68*, 601-604.
- [19] N. Della Ca', P. Bottarelli, A. Dibenedetto, M. Aresta, B. Gabriele, G. Salerno, M. Costa, *J. Catal.* **2011**, *282*, 120-127.
- [20] The use of ionic liquids as non-conventional solvents in organic synthesis has recently gained increasing attention, due to the very important characteristics of these non-conventional solvents; they are stable, non-flammable, non-volatile, recyclable, and in some cases may even promote organic reactions. For recent books and reviews, see: a) Z. S. Qureshi, K. M. Deshmukh, B. M. Bhanage, *Clean Technol. Environ. Policy* **2014**, *16*, 1487-1513; b) S. J. Zhang, J. Sun, X. Zhang, J. Xin, Q. Miao, J. Wang, *Chem. Soc. Rev.* **2014**, *43*, 7838-7869; c) P. Zhang, T. Wu, B. Han, *Adv. Mater.* **2014**, *26*, 6810-6827; d) R. Skoda-Foeldes, *Molecules* **2014**, *19*, 8840-8884; e) T. Itoh, *J. Synth. Org. Chem. Jpn.* **2014**, *72*, 518-528; f) J. Nowicki, M. Muszynski, *Curr. Org. Chem.* **2014**, *18*, 2797-2807; g) G. Chatel, D. R. MacFarlane, *Chem. Soc. Rev.* **2014**, *43*, 8132-8149; h) M. M. Cecchini, C. Charnay, F. De Angelis, F. Lamaty, J. Martinez, E. Colacino, *ChemSusChem* **2014**, *7*, 45-65; i) T. L. Greaves, C. J. Drummond, *Chem. Soc. Rev.* **2013**, *42*, 1096-1120; j) A. Rehman, X. Zeng, *Acc. Chem. Res.* **2012**, *45*, 1667-1677; k) D. D. Patel, *Chem. Rec.* **2012**, *12*, 329-355; l) W.-L. Wong, K.-Y. Wong, *Can. J. Chem.* **2012**, *90*, 1-16; m) T. Payagala, D. W. Armstrong, *Chirality* **2012**, *24*, 17-53; n) *Ionic Liquids in Biotransformations and Organocatalysis: Solvents and Beyond* (Ed. P. Dominguez de Maria, Wiley-VCH, Weinheim, Germany, **2012**); o) J. Dupont, *Acc. Chem. Res.* **2011**, *44*, 1223-1231; p) J. P. Hallett, T. Welton, *Chem. Rev.* **2011**, *111*, 3508-3576; q) C. D. Hubbard, P. Illner, R. van Eldik, *Chem. Soc. Rev.* **2011**, *40*, 272-290; r) Q. Zhang, S. Zhang, Y. Deng, *Green Chem.* **2011**, *13*, 2619-2637; s) *Ionic Liquids: Theory, Properties, New Approaches* (Ed. A. Kokorin, InTech, Rijeka, Croatia, **2011**); t) *Ionic Liquids: Applications and Perspectives* (Ed. A. Kokorin, InTech, Rijeka, Croatia, **2011**); u) C. Chiappe, S. Rajamani, *Eur. J. Org. Chem.* **2011**, *28*, 5517-5539; w) *Ionic Liquids in Synthesis* (Ed. P. Wasserscheid, Wiley-VCH, Weinheim, Germany, **2006**).
- [21] B. Gabriele, M. Costa, G. Salerno, G. P. Chiusoli, *J. Chem. Soc., Perkin Trans. 1* **1994**, 83-87.
- [22] Following a referee's suggestion, in order to verify the effect of the ionic liquid on the reactivity of secondary amines, we carried out some experiments in DME-BmimBF₄ mixtures with different v/v ratios (a: 95/5; b: 75/25; c: 50/50; d: 25/75; and e: 5/95) using morpholine **8a** as substrate, under the same conditions reported in Table 4, entry 1, run 1. A consistent trend was found, the yield in the corresponding oxamide **9a** increasing with the augment of the relative amount of the IL: 18% (mixture a); 28% (b); 34% (c); 48% (d); and 65% (e). From these data, we may conclude that the IL does not participate as a catalyst for the process; rather, its high dielectric constant may be responsible for the co-(activation) of the substrate and/or the Pd catalyst.

Entry for the Table of Contents (Please choose one layout)

Layout 1:

FULL PAPER

The first general recyclable catalytic oxidative carbonylation of amines, β -amino alcohols and 2-aminophenols to ureas, oxamides, oxazolidinones, and benzoxazolones is reported.



Raffaella Mancuso,^{*,[a,b]} Dnyaneshwar S. Raut,^[a] Nicola Della Ca,^[c] Francesco Fini,^[d] Carla Carfagna,^[e] and Bartolo Gabriele^{*,[a]}

Page No. – Page No.

Recyclable Catalytic Oxidative Carbonylation of Amines, β -Amino alcohols, and 2-Aminophenols to Ureas, Oxamides, 2-Oxazolidinones, and Benzoxazolones